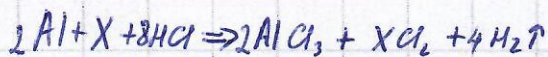


№1 Дано:



$$Al + X = 22,80 \text{ г}$$

$$(H_2)n = \frac{24,64}{22,4} = 1,1 \text{ моль}$$

$$V(H_2) = 24,64 \text{ л}$$

$$n(Al) = n(H_2) : 4 : 1 = 1 : 4 = 0,275 \text{ моль}$$

$$X > Al \text{ в } 1,25 \text{ р.}$$

$$m(Al) = 0,275 \cdot 27 = 7,425 \text{ г}$$

Найти:

$$m(Al + X) = 22,8 - 7,425 = 15,375$$

X - ?

$$m(X) = 15,375 : 0,275 = 55,9 \text{ г/моль}$$

 $w(X) = ?$

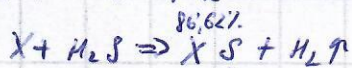
$$Fe = X$$

$$w(Fe) = \frac{15,375}{22,8} \cdot 100\% = 67,5\%$$

№2 Дано:



X - мет.



$$N(H_2) = 22,4 \cdot 1 = 22,4 \text{ л}$$

$$V(H_2S) = 22,4 \text{ л } (H_2)$$

$$100 - 86,62 = 13,38 \%$$

$$S - 13,38\%$$

$$X - 86,62\%$$

$$X = \frac{32 \cdot 86,62}{13,38} = 207,16 \text{ г/моль}$$

$$X = Pb$$

